

Using your Solubility Rules, write the formula of the following compounds and identify them as: **SE** = strong electrolyte, **WE** = weak electrolyte, **M** = molecular, or **P** = precipitate

Compound	Formula	Characteristic
1. Dinitrogen pentoxide	_____	SE WE M P
2. Barium sulfate	_____	SE WE M P
3. Mercury (I) bromide	_____	SE WE M P
4. Acetic acid	_____	SE WE M P
5. Hydroiodic acid	_____	SE WE M P
6. Carbon dioxide	_____	SE WE M P
7. Nitric acid	_____	SE WE M P
8. Sulfuric acid	_____	SE WE M P
9. Ammonium oxalate	_____	SE WE M P
10. Nickel (II) phosphate	_____	SE WE M P

### Net Ionic Equations (NIE)

For #1-3, write the NIE using the 3-step method. For #4-5, use your most efficient method to solve for the NIE. (3-step method *or your own technique*)

- Potassium phosphate with mercury (I) acetate → potassium acetate and mercury (I) phosphate

Molecular equation:

Total Ionic equation:

Net Ionic equation:

2. Sodium sulfide with iron (III) nitrate → sodium nitrate and iron (III) sulfide

Molecular equation:

Total Ionic equation:

Net Ionic equation:

3. Solutions of cesium hydroxide and oxalic acid are mixed.

Molecular equation:

Total Ionic equation:

Net Ionic equation:

4. Solutions of magnesium perchlorate and lead (II) nitrate are mixed.

5. Solutions of chromium (II) bromide with lithium oxalate are mixed.